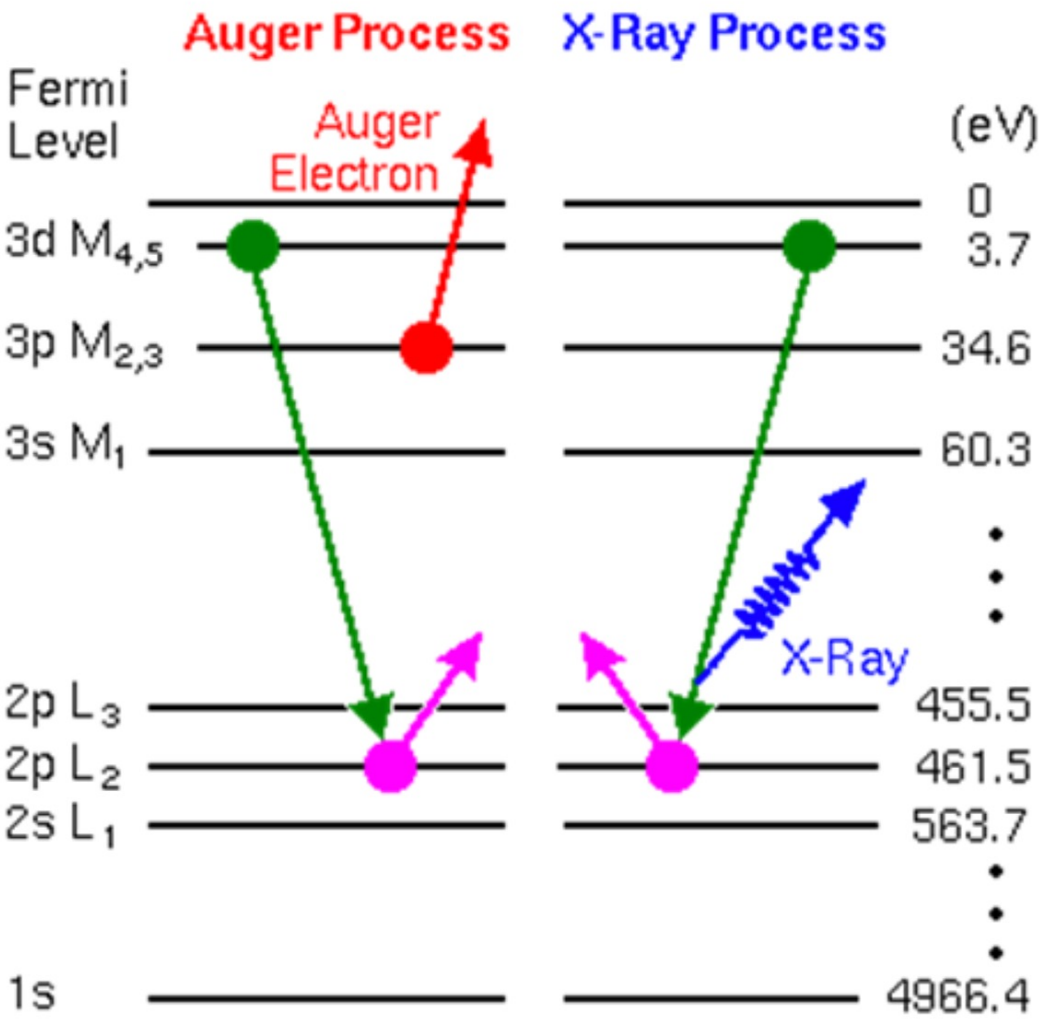
Energy of Auger electrons and of radiative emission

A core electron (violet arrow) of Ti atoms in a titanium film is excited by an external electron. The two possible relaxation processes are the Auger electron emission or the X-ray emission.

Calculate the energy of the emitted Auger electron or the emitted photon.
Where necessary, consider a work function = 5 eV.

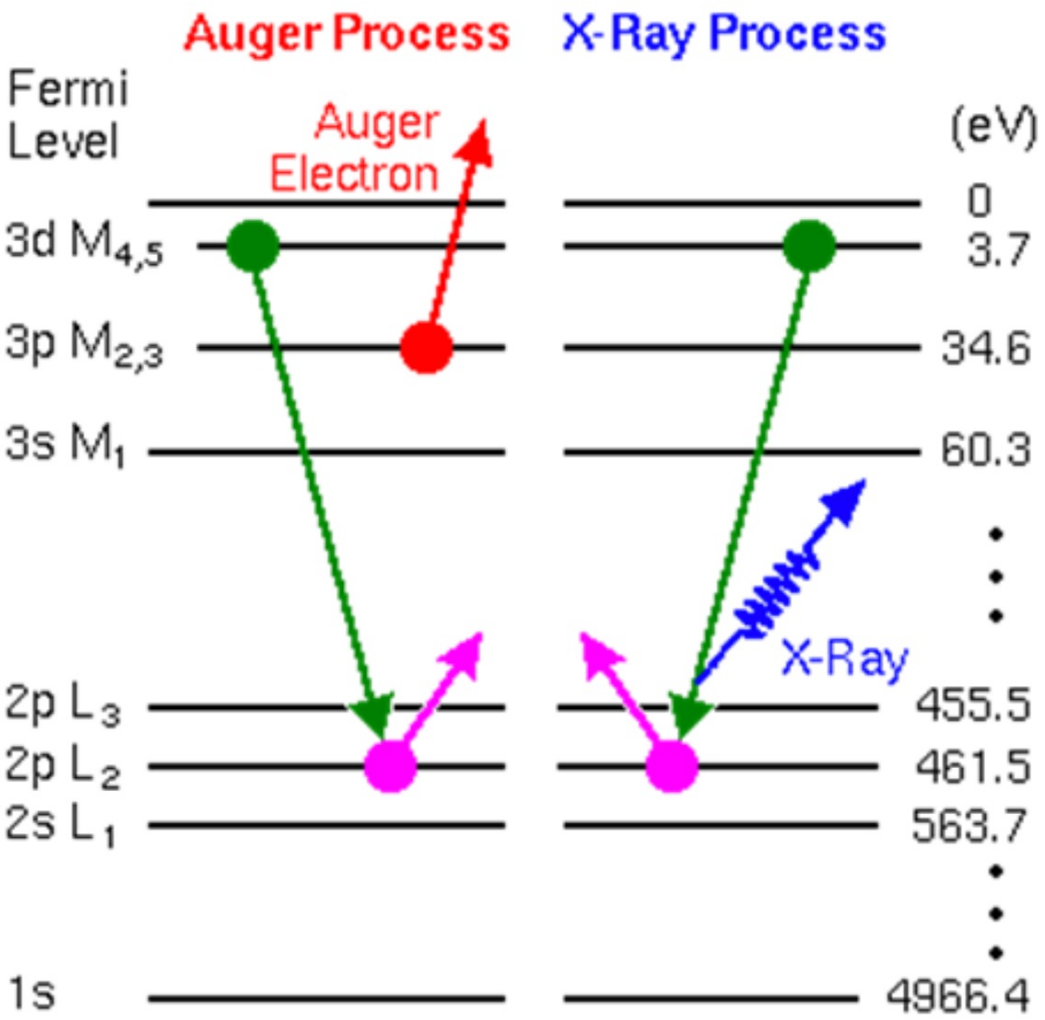


Solution: Energy of Auger electrons and of radiative emission

The Auger electron energy is

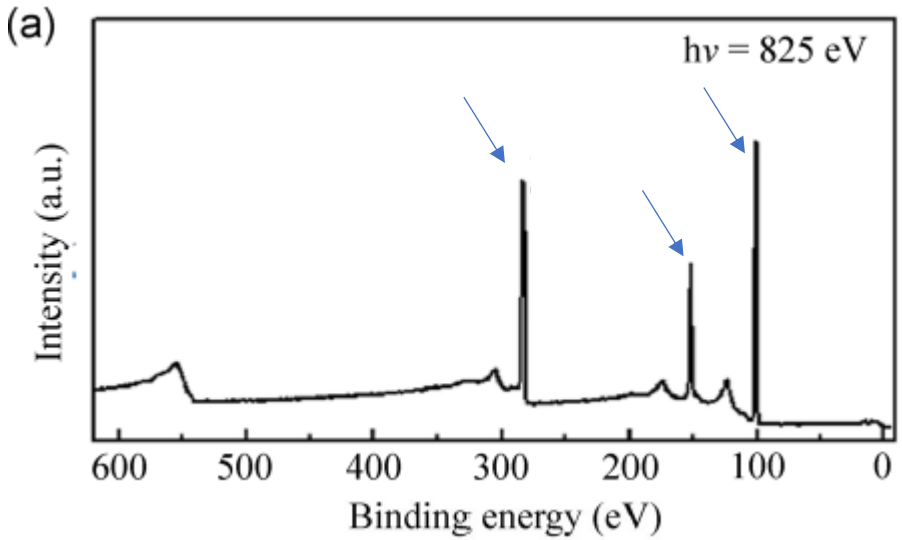
$E_{kin} = E_{L2} - E_{M4} - E_{M2} - \Phi \sim 418$ eV
LMM transition

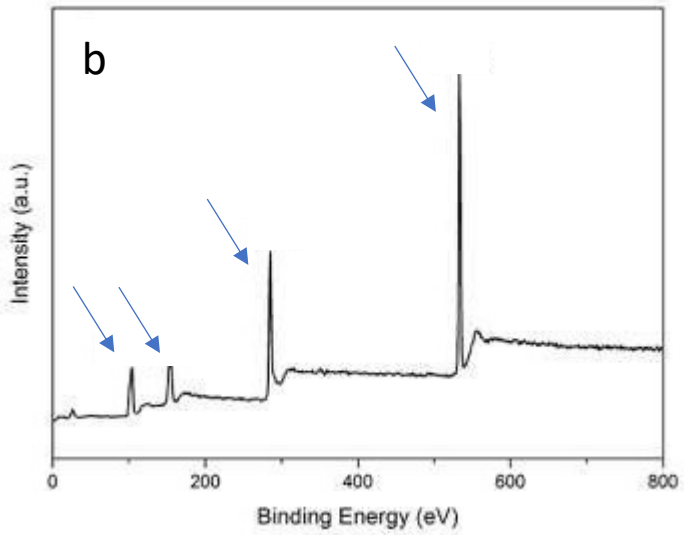
The X-ray photon energy is
$E_{hv} = E_{L2} - E_{M4} \sim 457.8$ eV

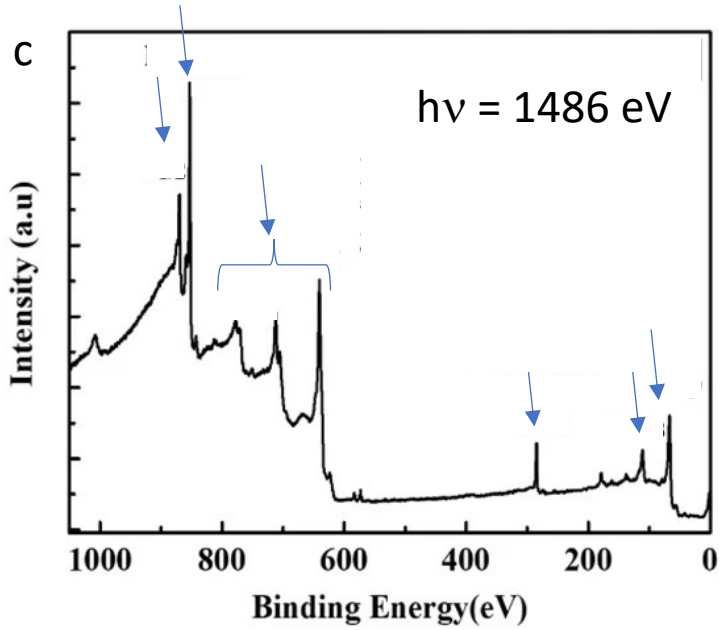


XPS spectra of graphene on different substrates

The XPS spectra of graphene on three unknown substrates are shown below.
Identify the major peaks marked with arrows and determine the three substrates
(Use the table of the electron binding energies for potential element candidates provided in the next page)







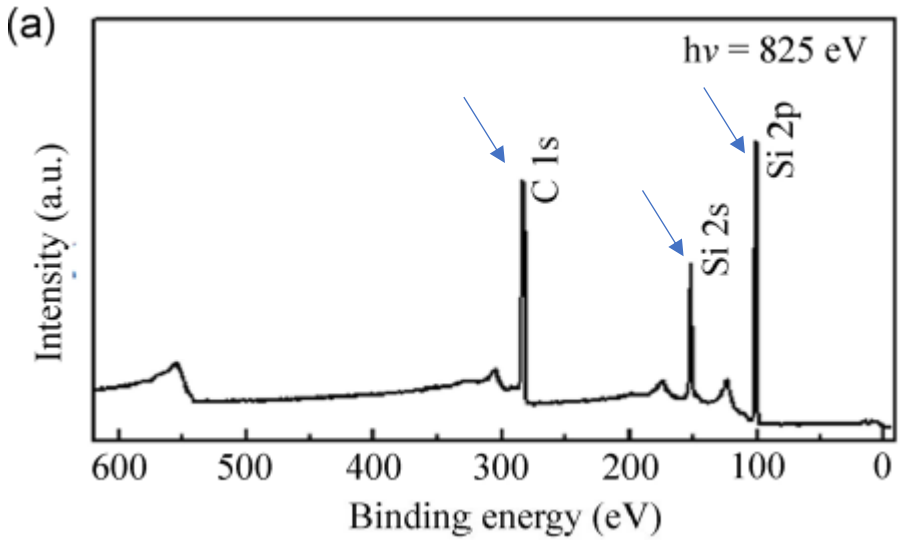
Sources: DOI:10.1039/C6RA05884A; DOI:10.1007/s12274-014-0584-y; doi.org/10.3390/nano11113112

Table 1-1. *Electron binding energies, in electron volts, for the elements in their natural forms.*

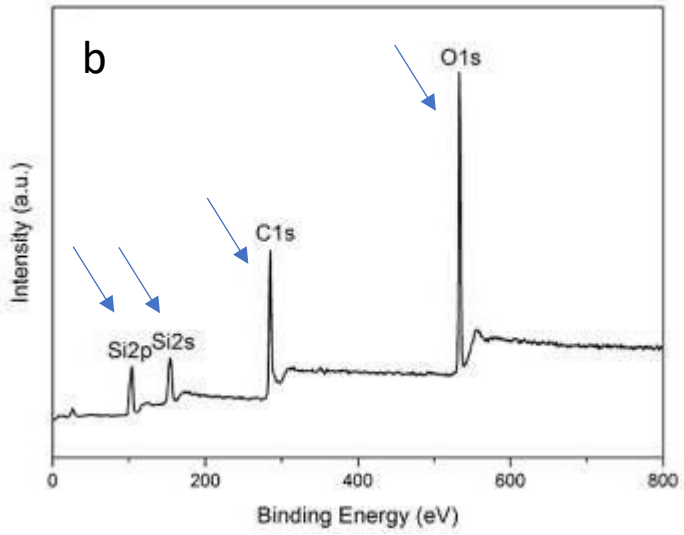
Element	K 1s	L_1 2s	L_2 $2p_{1/2}$	L_3 $2p_{3/2}$	M_1 3s	M_2 $3p_{1/2}$	M_3 $3p_{3/2}$	M_4 $3d_{3/2}$	M_5 $3d_{5/2}$	N_1 4s	N_2 $4p_{1/2}$	N_3 $4p_{3/2}$
1 H	13.6											
2 He	24.6*											
3 Li	54.7*											
4 Be	111.5*											
5 B	188*											
6 C	284.2*											
7 N	409.9*	37.3*										
8 O	543.1*	41.6*										
9 F	696.7*											
10 Ne	870.2*	48.5*	21.7*	21.6*								
11 Na	1070.8†	63.5†	30.65	30.81								
12 Mg	1303.0†	88.7	49.78	49.50								
13 Al	1559.6	117.8	72.95	72.55								
14 Si	1839	149.7*b	99.82	99.42								
15 P	2145.5	189*	136*	135*								
16 S	2472	230.9	163.6*	162.5*								
17 Cl	2822.4	270*	202*	200*								
18 Ar	3205.9*	326.3*	250.6†	248.4*	29.3*	15.9*	15.7*					
19 K	3608.4*	378.6*	297.3*	294.6*	34.8*	18.3*	18.3*					
20 Ca	4038.5*	438.4†	349.7†	346.2†	44.3 †	25.4†	25.4†					
21 Sc	4492	498.0*	403.6*	398.7*	51.1*	28.3*	28.3*					
22 Ti	4966	560.9†	460.2†	453.8†	58.7†	32.6†	32.6†					
23 V	5465	626.7†	519.8†	512.1†	66.3†	37.2†	37.2†					
24 Cr	5989	696.0†	583.8†	574.1†	74.1†	42.2†	42.2†					
25 Mn	6539	769.1†	649.9†	638.7†	82.3†	47.2†	47.2†					
26 Fe	7112	844.6†	719.9†	706.8†	91.3†	52.7†	52.7†					
27 Co	7709	925.1†	793.2†	778.1†	101.0†	58.9†	59.9†					
28 Ni	8333	1008.6†	870.0†	852.7†	110.8†	68.0†	66.2†					
29 Cu	8979	1096.7†	952.3†	932.7	122.5†	77.3†	75.1†					
30 Zn	9659	1196.2*	1044.9*	1021.8*	139.8*	91.4*	88.6*	10.2*	10.1*			
31 Ga	10367	1299.0*b	1143.2†	1116.4†	159.5†	103.5†	100.0†	18.7†	18.7†			
32 Ge	11103	1414.6*b	1248.1*b	1217.0*b	180.1*	124.9*	120.8*	29.8	29.2			
33 As	11867	1527.0*b	1359.1*b	1323.6*b	204.7*	146.2*	141.2*	41.7*	41.7*			
34 Se	12658	1652.0*b	1474.3*b	1433.9*b	229.6*	166.5*	160.7*	55.5*	54.6*			
35 Br	13474	1782*	1596*	1550*	257*	189*	182*	70*	69*			
36 Kr	14326	1921	1730.9*	1678.4*	292.8*	222.2*	214.4	95.0*	93.8*	27.5*	14.1*	14.1*
37 Rb	15200	2065	1864	1804	326.7*	248.7*	239.1*	113.0*	112*	30.5*	16.3*	15.3 *
38 Sr	16105	2216	2007	1940	358.7†	280.3†	270.0†	136.0†	134.2†	38.9†	21.3	20.1†
39 Y	17038	2373	2156	2080	392.0*b	310.6*	298.8*	157.7†	155.8†	43.8*	24.4*	23.1*
40 Zr	17998	2532	2307	2223	430.3†	343.5†	329.8†	181.1†	178.8†	50.6†	28.5†	27.1†
41 Nb	18986	2698	2465	2371	466.6†	376.1†	360.6†	205.0†	202.3†	56.4†	32.6†	30.8†
42 Mo	20000	2866	2625	2520	506.3†	411.6†	394.0†	231.1†	227.9†	63.2†	37.6†	35.5†
43 Tc	21044	3043	2793	2677	544*	447.6	417.7	257.6	253.9*	69.5*	42.3*	39.9*
44 Ru	22117	3224	2967	2838	586.1*	483.5†	461.4†	284.2†	280.0†	75.0†	46.3†	43.2†
45 Rh	23220	3412	3146	3004	628.1†	521.3†	496.5†	311.9†	307.2†	81.4*b	50.5†	47.3†
46 Pd	24350	3604	3330	3173	671.6†	559.9†	532.3†	340.5†	335.2†	87.1*b	55.7†a	50.9†
47 Ag	25514	3806	3524	3351	719.0†	603.8†	573.0†	374.0†	368.3	97.0†	63.7†	58.3†

Solution: XPS spectra of graphene on different substrates

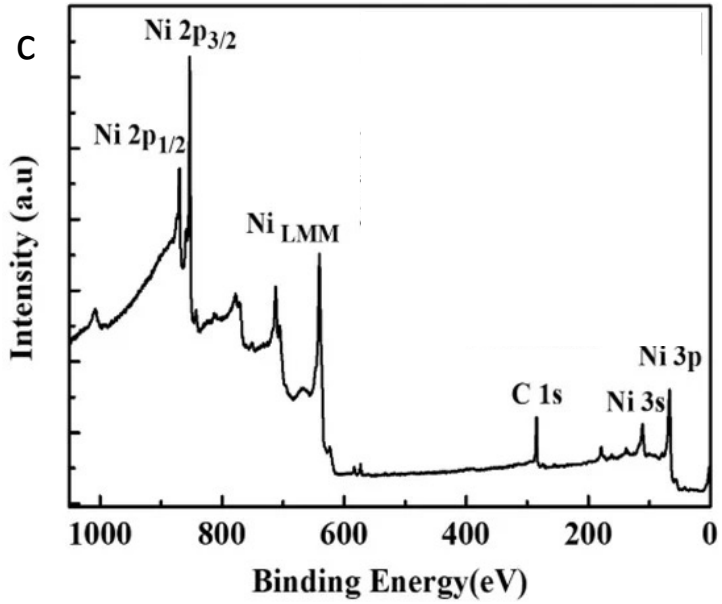
Graphene on Si



Graphene on SiO_2



Graphene on Ni



UPS spectra

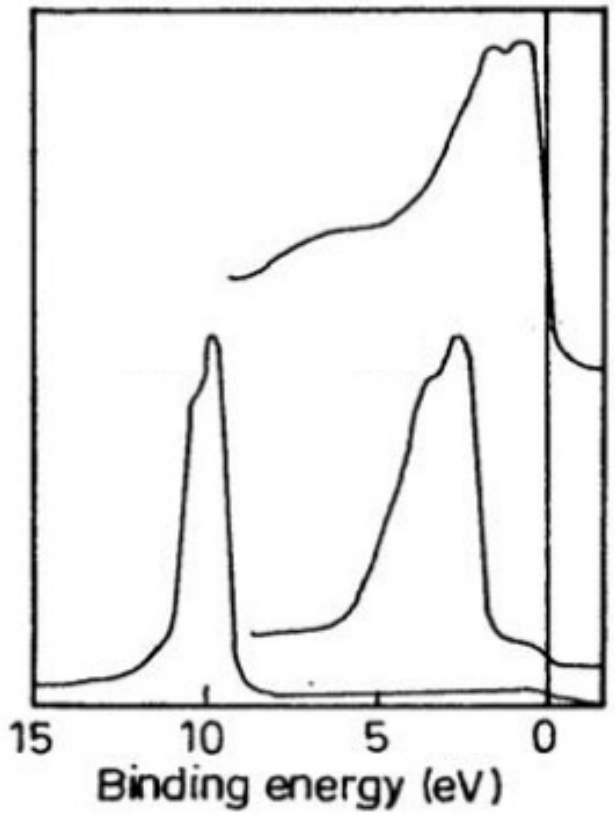
Consider the three UPS spectra shown in the figure, where one has been vertically shifted for clarity.

Using electron configuration given for the three elements (gas phase), identify the spectrum of the solid formed by each element. Justify you answers.

Ni (0): $[Ar]3d^8 4s^2$

Cu(0): $[Ar]3d^{10}4s^1$

Zn (0): $[Ar]3d^{10}4s^2$



Solution: UPS spectra

From the filling of the 3d and 4s states, we can deduce that:

- the spectrum with the peak at E_{bind} = 10 eV corresponds to Zn, since the 3d states are fully occupied;

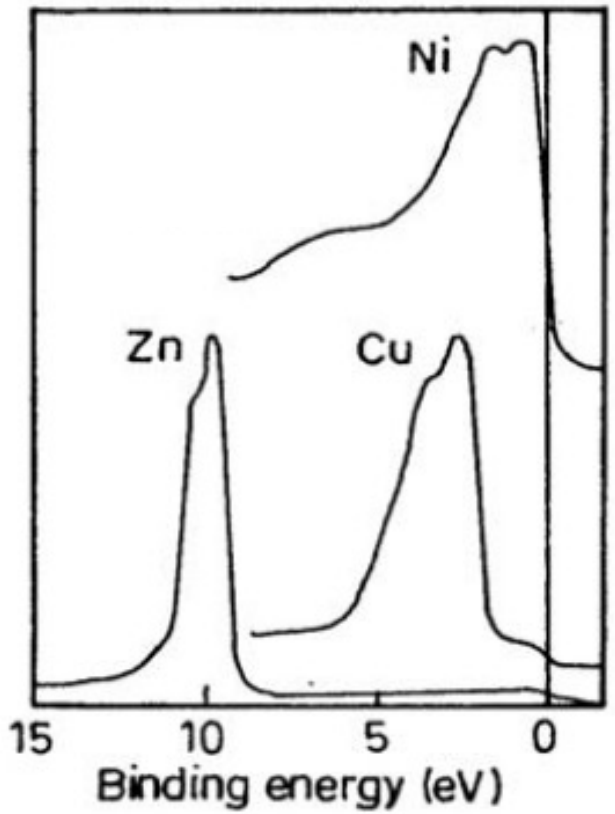
- the spectrum with the peak at E_{bind} = 2-5 eV corresponds to Cu, since the 3d states are fully occupied but we expect the 4s states to be partially occupied;

- the spectrum with the peak at E_{bind} = 0-3 eV corresponds to Ni, since the 3d states are not fully occupied.

Ni (0): $[Ar]3d^8 4s^2$

Cu(0): $[Ar]3d^{10}4s^1$

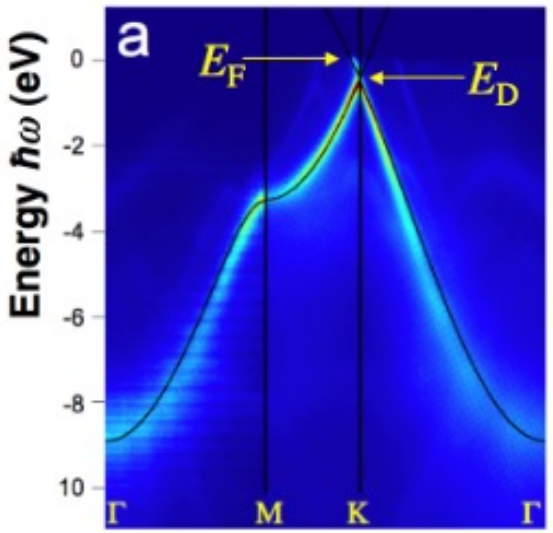
Zn (0): $[Ar]3d^{10}4s^2$

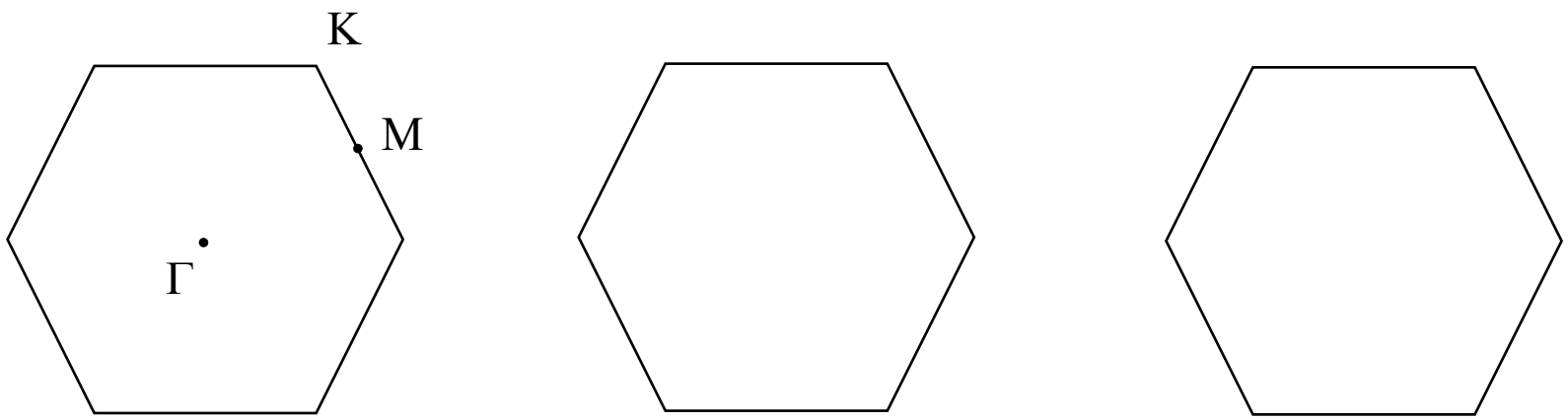


Graphene - Constant energy surfaces

The ARPES spectrum shows the π bands measured on graphene grown on SiC.

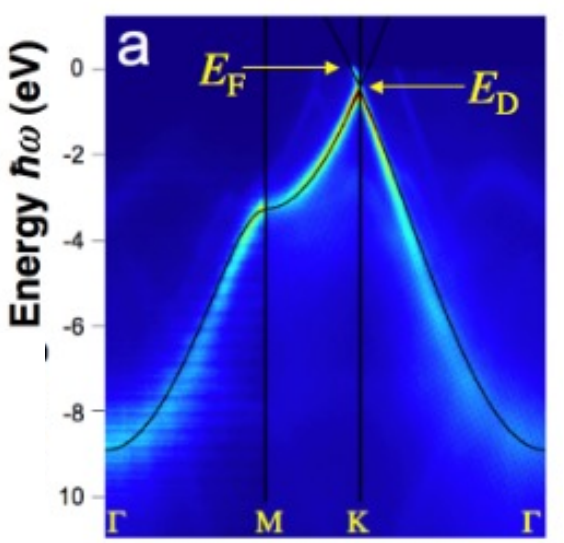
Plot the constant energy surface at $E = E_D$, at $E = E_F$ (Fermi surface), and at $E = -2$ eV below E_F

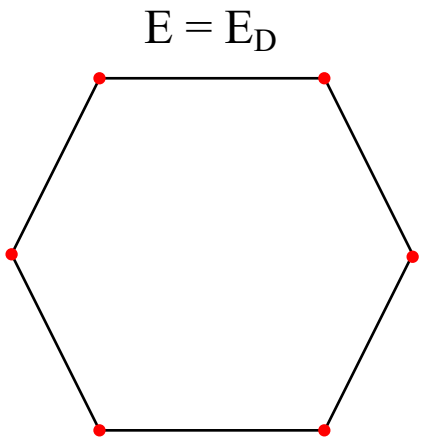


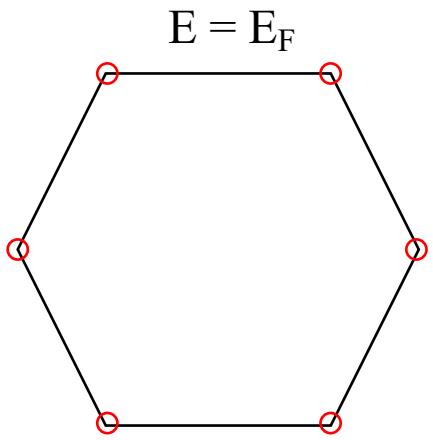


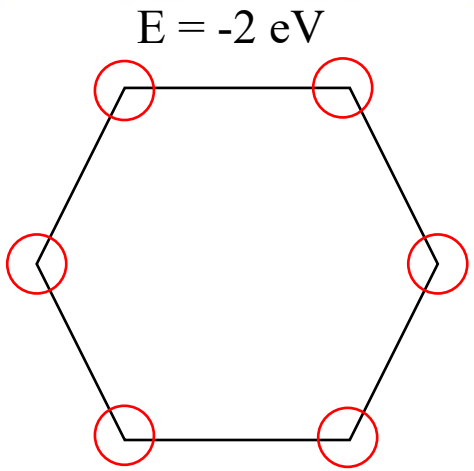
Adapted from *Nature Materials* **6,** 183–191 (2007)

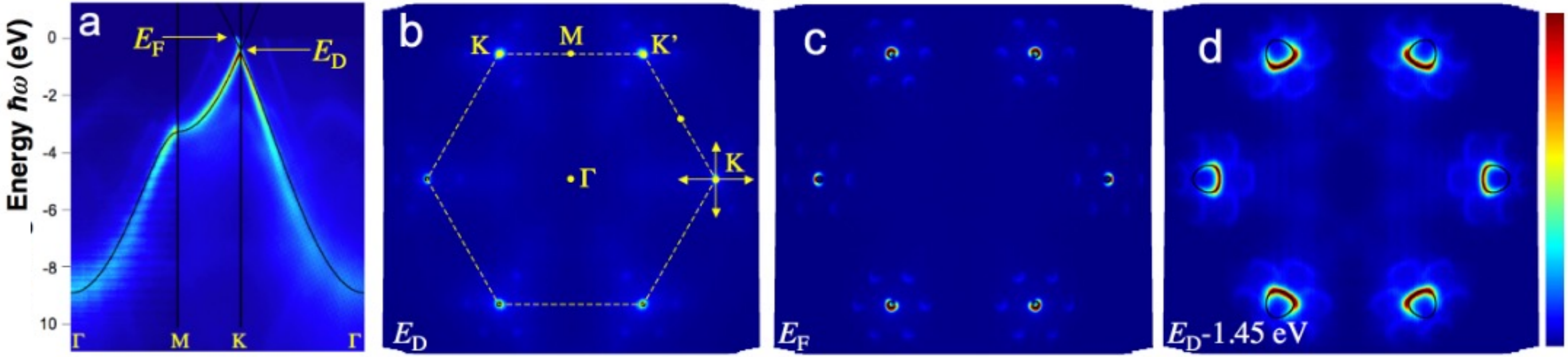
Solution: Graphene – Constant energy surfaces











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